

## OCEAN pH LEVELS INQUIRY LABS – SHORT

# The pH of Seawater

The surface of the ocean has an average pH of 8.1, that is approximately 0.1 units lower than before the industrial revolution started. The decrease in pH of seawater is directly related to the increase in carbon dioxide emissions to the atmosphere, mostly from human activities involving the burning of fossil fuels. In this laboratory, you will investigate the equilibrium between water, carbon dioxide, bicarbonate, and carbonate species, and its role in determining the pH levels of ocean water.

## Focus on Science Practices

**SEP 2** Developing and Using Models**SEP 4** Analyzing and Interpreting Data**SEP 6** Constructing Explanations and Designing Solutions

## Materials Per Group

- Acetic acid solution,  $\text{CH}_3\text{COOH}$ , 2 M, 30 mL
- Sodium bicarbonate,  $\text{NaHCO}_3$ , 3.5 g
- Sodium carbonate,  $\text{Na}_2\text{CO}_3$ , 0.5 g
- Sodium chloride,  $\text{NaCl}$ , 6 g
- Phenolphthalein indicator, 1% solution, 0.5 mL
- Water, distilled or deionized
- Beakers, 400 mL and 250 mL
- Beral-type pipet
- Clamp
- Heat-resistant gloves
- Hot plate
- Erlenmeyer flask for filtering, 250 mL
- Graduated cylinder, 50 mL and 100 mL
- pH meter
- pH test strips (optional)
- Plastic tubing, 2–3 feet long
- Ring stand
- Rubber stopper (for use with Erlenmeyer flask)
- Spatula
- Weighing dishes, 3

## Safety

*Sodium bicarbonate is slightly toxic by ingestion. Sodium carbonate is toxic by ingestion, and can cause skin and eye irritation. Acetic acid solution is toxic by ingestion and inhalation, and is corrosive to skin and eyes. Wear chemical splash goggles, chemical-resistant gloves, and a chemical-resistant apron when handling these*

*chemicals. Avoid exposure to eyes and skin, and clean up all spills promptly. Wash hands thoroughly with soap and water before leaving the laboratory.*

## Procedure

1. Prepare 100 mL of a 0.5 M sodium chloride solution that will resemble seawater, by dissolving the appropriate amount of NaCl in deionized/distilled water. Record your detailed calculations.
2. Use a graduated cylinder to measure and transfer 100 mL of the seawater solution prepared in Step 1 into a clean beaker. Place a pH meter in the solution and measure its pH.
3. Use the spatula to add a very small amount of sodium carbonate ( $\text{Na}_2\text{CO}_3$ ) to the solution to adjust its pH to make it as close as possible to 8.0. Use the tip of the spatula to add the sodium carbonate. If the pH goes beyond 9, add small amounts of sodium bicarbonate to lower the pH to about 8.
4. Add a few drops of phenolphthalein indicator to the solution, and record your observations in the data table.
5.  Attach the plastic tubing to the short glass tube on the Erlenmeyer flask. At first, it may be helpful to heat up one end of the plastic tubing by placing it in hot water to stretch it out a little. This should facilitate sliding it onto the glass tube on the Erlenmeyer flask.
6. Measure 20 mL of 2 M acetic acid solution using a graduated cylinder, and transfer into the Erlenmeyer flask.
7. Measure approximately 3.0 g of sodium bicarbonate in a weighing dish or paper.
8. Dip the loose end of the plastic tubing into the beaker containing the seawater solution with the universal indicator. Secure the flask containing the seawater solution by placing it into a larger beaker/flask. Secure the Erlenmeyer flask containing the acetic acid solution using a clamp firmly attached to a ring stand.
9. Transfer the sodium bicarbonate solid into the Erlenmeyer flask containing the acetic acid solution, and quickly put on the rubber stopper to close the flask.

10.  Loosen the Erlenmeyer flask from the clamp, and begin to stir the contents gently. Gaseous carbon dioxide will be produced in the flask, and it will flow into the seawater solution in the beaker through the plastic tubing.
11. Monitor the pH and color of the seawater solution. Write your observations in the data table.
12. Continue to measure the pH of the seawater solution until it becomes stable. Record this pH value as the final pH of the solution.
13. Remove the pH meter from the seawater solution, rinse it with distilled/deionized water, and store it as indicated by your instructor.
14. Dispose of the contents of the Erlenmeyer flask and the seawater solution as indicated by your instructor.

<b>Data Table—The pH of Seawater</b>	
<b>Parameter</b>	<b>Observations (color, appearance, etc.)</b>

pH before treatment with CO <sub>2</sub>		
[H <sup>+</sup> ] before treatment with CO <sub>2</sub> (moles/liter)		
pH after treatment with CO <sub>2</sub>		
[H <sup>+</sup> ] after treatment with CO <sub>2</sub> (moles/liter)		
Acidity percent change (%)		

### Analyze and Interpret Data

- 1. SEP Use Models** Write a balanced chemical equation to represent the reaction(s) that takes place in the Erlenmeyer flask.
  
- 2. SEP Use Models** The dissolution of carbon dioxide in water produces carbonic acid (H<sub>2</sub>CO<sub>3</sub>) at first, which then dissociates into bicarbonate (HCO<sub>3</sub><sup>-</sup>) and hydrogen ions (H<sup>+</sup>). Write a balanced equation for this reaction (as an equilibrium).
  
- 3. SEP Use Math** Using the equation that defines pH, calculate the concentration of hydrogen ions ([H<sup>+</sup>] in moles/liter) before and after treating the seawater solution with carbon dioxide.

**4. SEP Use Math** Based on the results of your calculations in question 3, determine the percent change in acidity, that is the percent change in  $[H^+]$ , for the seawater solution upon treatment with carbon dioxide.

**5. SEP Construct an Explanation** Explain what happens to the seawater solution when the gas generated in the Erlenmeyer flask bubbles into it. How does the pH and the color of the solution change, and why?

**6. SEP Engage in Argument** Surface ocean water has a pH of approximately 8.1,

and it contains dissolved carbon dioxide in equilibrium with water, bicarbonate ions ( $\text{HCO}_3^-$ ), and, to a lesser extent, carbonate ions ( $\text{CO}_3^{2-}$ ). Based on your results from this investigation, how would an increase in atmospheric carbon dioxide influence the pH and acidity of seawater and why?